Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/13**

Paper 1 Multiple Choice **May/June 2025**

1 hour 15 minutes

You must answer on the multiple choice answer sheet.

1192205922

You will need: Multiple choice answer sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.
- Important values, constants and standards are printed in the question paper.

This document has **20** pages. Any blank pages are indicated.

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[Turn over

1 Which isolated gaseous atom has a total of five electrons occupying spherically shaped orbitals?

A sodium

B fluorine

C boron

D potassium

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2 Diagram 1 shows the reaction pathway for a reaction.

Diagram 2 shows a graph of concentration of reactant against time for the same reaction.

diagram 1

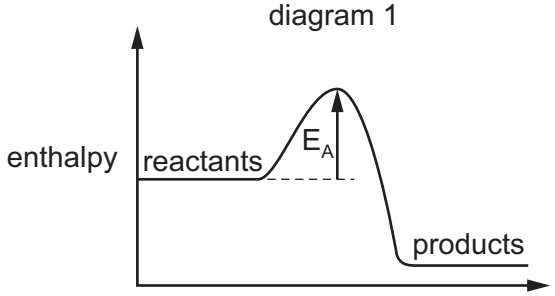
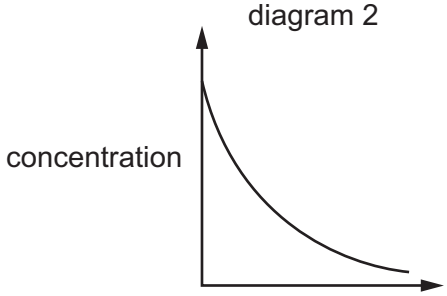
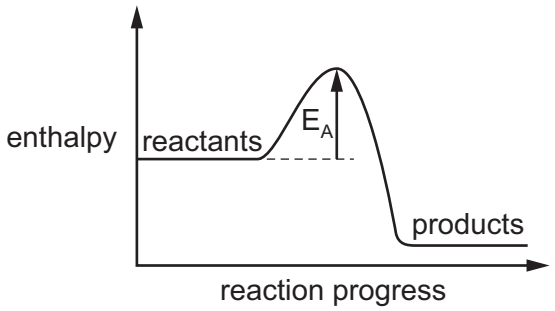


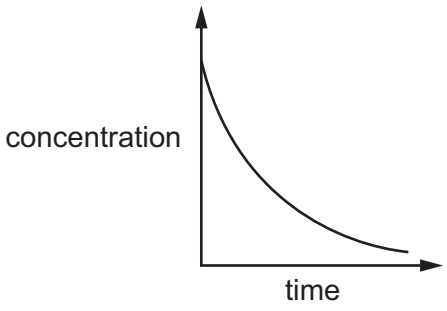
diagram 2



Which diagrams, drawn to the same scale, show the effect of increased temperature on this reaction?

A





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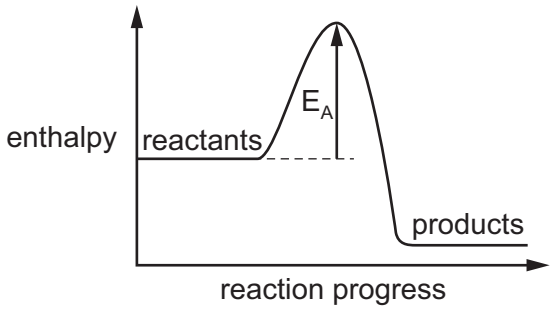

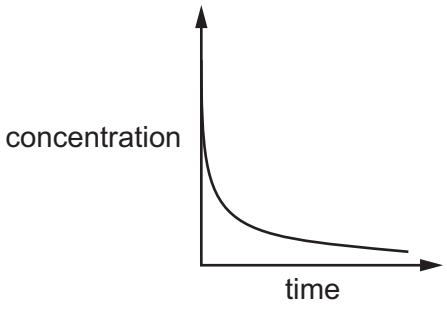

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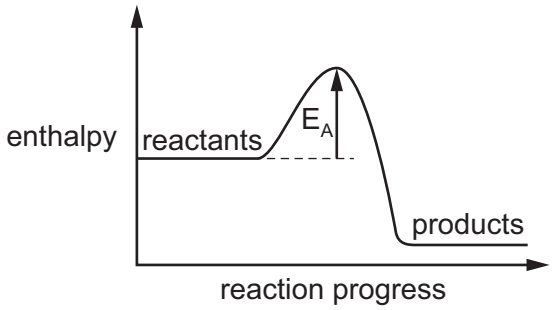

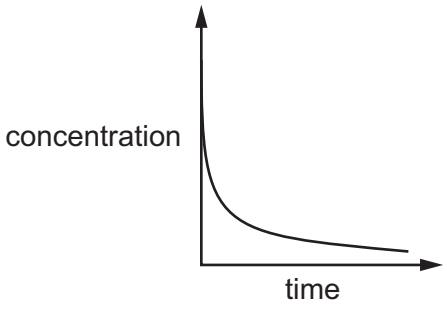

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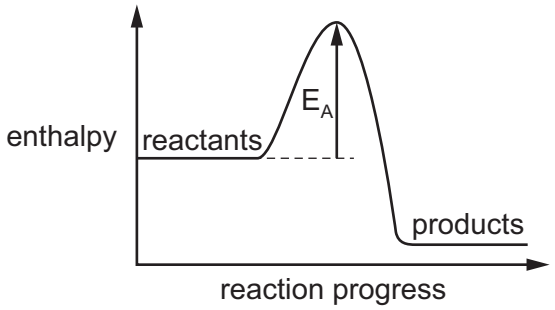
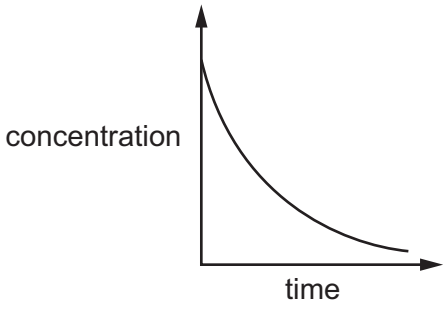
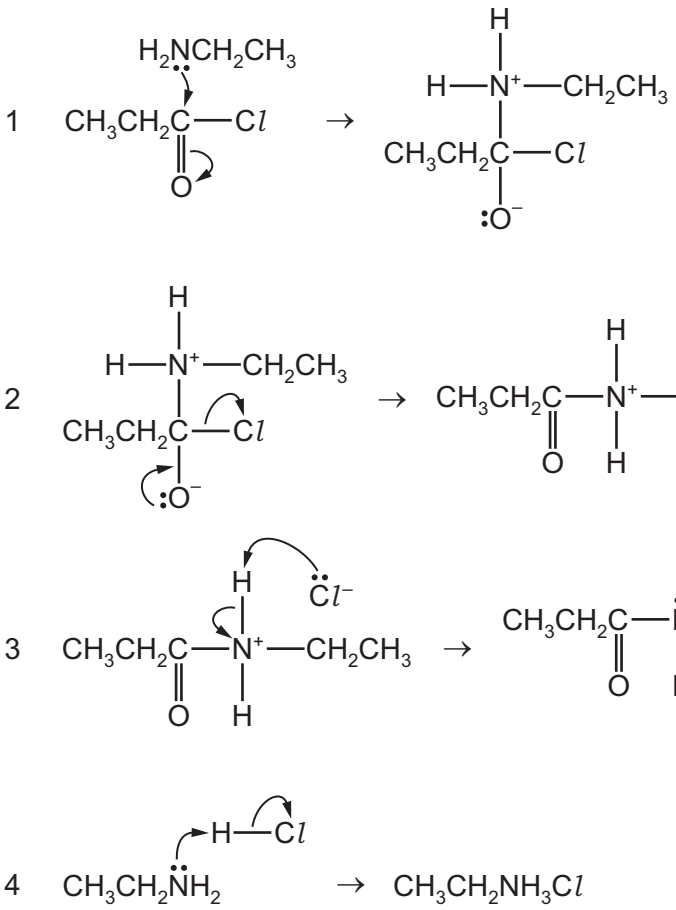




3 The equation shows the overall reaction between an amine and an acid chloride.

$$2CH_3CH_2NH_2 + CH_3CH_2COCl \rightarrow CH_3CH_2CONHCH_2CH_3 + CH_3CH_2NH_3Cl$$

The four steps in the mechanism are as follows.



Which steps are Brønsted–Lowry acid–base reactions?

A 1 and 3 **B** 3 and 4 **C** 3 only **D** 4 only

4 Chromium is present in compound X.

- Two moles of compound X react with exactly 3 moles of silicon.
- The only products of this reaction are 4 moles of chromium and 3 moles of a silicon compound in which the oxidation state of the silicon is +4.
- Chromium and silicon are the only elements that change their oxidation states in this reaction.

What could be the identity of compound X?

A Cr_2O_3 **B** Cr_2H_6 **C** Cr_2H_4 **D** Cr_2O_4

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5 The reversible reaction shown is in equilibrium at a temperature of 450 °C.

$$H_2 + I_2 \rightleftharpoons 2HI \qquad \Delta H = -9.5\,\text{kJ mol}^{-1}$$

The table shows the equilibrium concentrations in the reaction mixture.

H_2	I_2	HI
$1.02\,\text{mol dm}^{-3}$	$0.02\,\text{mol dm}^{-3}$	$0.98\,\text{mol dm}^{-3}$

The temperature of the reaction mixture is increased at constant volume.

The concentration of one of the components **falls** to $0.80\,\text{mol dm}^{-3}$ at equilibrium under these conditions.

What is the concentration of H_2 in the new equilibrium mixture?

A $0.80\,\text{mol dm}^{-3}$

B $0.93\,\text{mol dm}^{-3}$

C $1.11\,\text{mol dm}^{-3}$

D $1.20\,\text{mol dm}^{-3}$

6 The first stage in the industrial production of nitric acid from ammonia can be represented by the following equation.

$$4NH_3(g) + 5O_2(g) \rightleftharpoons 4NO(g) + 6H_2O(g)$$

Using the following standard enthalpy change of formation data, what is the value of the standard enthalpy change, $\Delta H^{\ominus}$, for this reaction?

compound	$\Delta H_f^{\ominus}$ / kJ mol^{-1}
$NH_3(g)$	–46.1
$NO(g)$	+90.3
$H_2O(g)$	–241.8

A $+905.2\,\text{kJ mol}^{-1}$

B $-105.4\,\text{kJ mol}^{-1}$

C $-905.2\,\text{kJ mol}^{-1}$

D $-1274.0\,\text{kJ mol}^{-1}$

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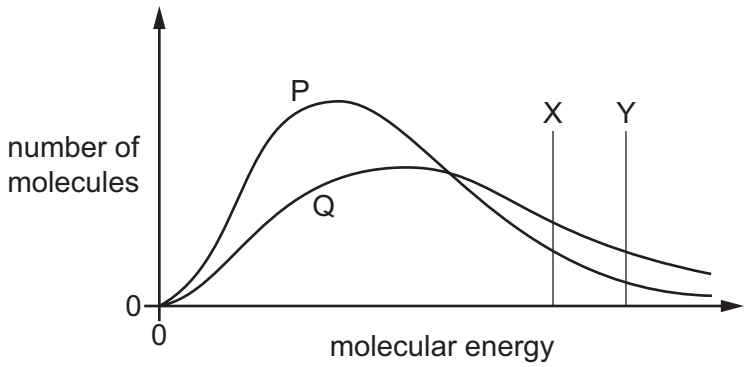
7 The gaseous compound Z decomposes on heating.

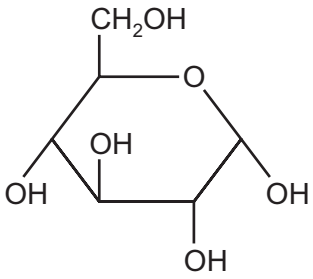
In the diagram, Boltzmann distributions for Z at two different temperatures, P and Q, are shown. The lines X and Y indicate activation energies for the decomposition of Z with and without a catalyst.



Which curve and which line describe the decomposition of Z at a higher temperature and with a catalyst present?

	higher temperature	catalyst present
A	P	X
B	P	Y
C	Q	X
D	Q	Y

8 The structure of a molecule found in food is shown.



Four statements about the molecule are listed.

1 This molecule contains only σ bonds.

2 This molecule contains both σ and π bonds.

3 The six-membered ring is planar.

4 The six-membered ring is non-planar.

Which two statements are correct?

A 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 2 and 4

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9 In which substance are covalent bonds broken as it melts?

A silicon(IV) oxide

B ice

C iodine

D ethanol

10 Why is the second ionisation energy of sodium larger than the second ionisation energy of magnesium?

A The attraction between the nucleus and the outer electron is greater in Na^+ than in Mg^+.

B The nuclear charge of Na^+ is greater than that of Mg^+.

C The outer electron of Na^+ is more shielded than the outer electron of Mg^+.

D The outer electron of Na is in the same orbital as the outer electron of Mg.

11 BF_3 reacts with NH_3 to form a single compound with a simple molecular structure.

$$BF_3 + NH_3 \rightarrow F_3B–NH_3$$

Which row describes how the bond angles change during the reaction?

	F–B–F bond angle	H–N–H bond angle
A	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

12 The equation for the reaction of nitrogen with hydrogen is shown.

$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \qquad \Delta H_r^{\ominus} = -92.2\,kJ\,mol^{-1}$$

Which statement is correct?

A $\Delta H_r^{\ominus}$ is measured at 298 °C.

B $\Delta H_r^{\ominus}$ is measured at 101 kPa.

C $\Delta H_r^{\ominus}$ represents the standard enthalpy change for the formation of ammonia gas.

D $\Delta H_r^{\ominus}$ represents the enthalpy change when 1.0 mol of $N_2(g)$ reacts with 1.0 mol of $H_2(g)$.

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13 Hydrogen peroxide, H_2O_2, decomposes into water and oxygen when a suitable catalyst is added.

$20.0\,cm^3$ of aqueous hydrogen peroxide decomposes to produce $600\,cm^3$ of oxygen at room conditions.

What is the concentration of the aqueous hydrogen peroxide?

A $5.00\,mol\,dm^{-3}$

B $2.50\,mol\,dm^{-3}$

C $1.25\,mol\,dm^{-3}$

D $0.625\,mol\,dm^{-3}$

14 When 0.15 g of an organic compound is vaporised, it occupies a volume of $65.0\,cm^3$ at 405 K and 1.00×10^5 Pa.

Using the expression $pV = nRT$, which expression should be used to calculate the relative molecular mass, M_r, of the compound?

A $\dfrac{0.15 \times 65 \times 10^{-6} \times 1 \times 10^5}{8.31 \times 405}$

B $\dfrac{0.15 \times 8.31 \times 405}{1 \times 10^5 \times 65 \times 10^{-3}}$

C $\dfrac{0.15 \times 65 \times 10^{-3} \times 1 \times 10^5}{8.31 \times 405}$

D $\dfrac{0.15 \times 8.31 \times 405}{1 \times 10^5 \times 65 \times 10^{-6}}$

15 A washing powder contains sodium hydrogencarbonate, $NaHCO_3$, as one of the ingredients.

In a titration, a solution containing 1.00 g of this washing powder requires $7.15\,cm^3$ of $0.100\,mol\,dm^{-3}$ sulfuric acid for complete reaction. The sodium hydrogencarbonate is the only ingredient that reacts with the acid.

What is the percentage by mass of sodium hydrogencarbonate in the washing powder?

A 3.0% **B** 6.0% **C** 12.0% **D** 24.0%

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16 The flow chart shows some reactions of nitrogen compounds.

$$\text{gas 1} \xrightarrow{\text{acid}} \text{salt 1} \xrightarrow{\text{alkali}} \text{salt 2 + other products}$$

Which row identifies gas 1 and salt 2?

	gas 1	salt 2
A	ammonia	ammonium nitrate
B	ammonia	sodium nitrate
C	nitrogen dioxide	ammonium nitrate
D	nitrogen dioxide	sodium nitrate

17 Mixing aqueous silver nitrate and aqueous sodium chloride produces a precipitate.

Addition of which reagent to the mixture gives a colourless solution?

A aqueous ammonia

B aqueous potassium iodide

C dilute hydrochloric acid

D dilute nitric acid

18 Which statement is correct?

A Barium oxide reacts with water at room temperature to form barium hydroxide and hydrogen.

B Calcium oxide does **not** react with water at room temperature.

C Magnesium hydroxide reacts with steam to form magnesium oxide and hydrogen.

D Strontium oxide reacts with water at room temperature to form strontium hydroxide only.

19 Which statement about the properties of halogens and hydrogen halides is correct?

A A chloride ion is a stronger reducing agent than an iodide ion.

B Chlorine is more volatile than bromine because chlorine has stronger intermolecular forces.

C Hydrogen bromide is more thermally stable than hydrogen iodide because it has a stronger covalent bond.

D Iodine is less reactive than bromine because iodine has weaker covalent bonds.

20 Three statements about the chemical periodicity of Period 3 oxides are listed.

1 The maximum oxidation state of the Period 3 elements in their oxides increases from sodium to phosphorus, then decreases from phosphorus to sulfur.

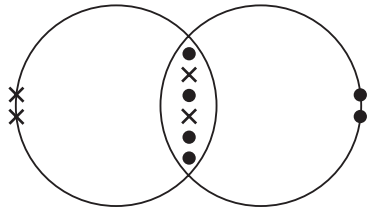
2 The oxides from sodium to aluminium dissolve in water without hydrolysis; from silicon to sulfur they are hydrolysed by water.

3 The structure and bonding changes from giant ionic to giant covalent to simple molecular.

Which statements are correct?

A 1, 2 and 3 **B** 1 only **C** 2 and 3 only **D** 3 only

21 Which dot-and-cross diagrams represent molecules that can react with unburned hydrocarbons to form peroxyacetyl nitrate, PAN, a component of photochemical smog?



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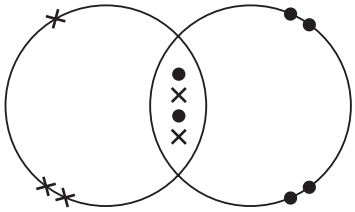

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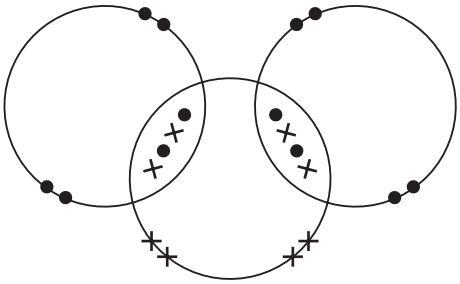

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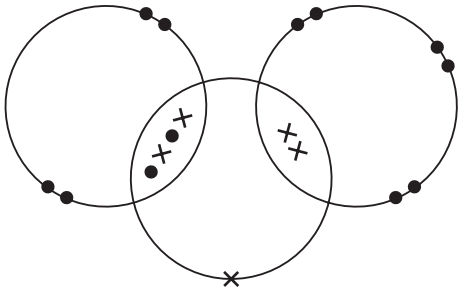

4

A 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 2 and 4

22 Which set of three elements contains a single element that has **both** the highest melting point **and** the lowest electrical conductivity of the three elements in the set?

A magnesium, aluminium and silicon

B aluminium, silicon and phosphorus

C sodium, magnesium and aluminium

D silicon, phosphorus and chlorine

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23 In which row do the particles increase in size?

	smallest →		largest
A	N	O	F
B	N^{3-}	O^{2-}	F^{-}
C	Na^{+}	Mg^{2+}	Al^{3+}
D	Na^{+}	Ne	F^{-}

24 W, X, Y and Z are Group 2 elements barium, calcium, magnesium and strontium but not in that order.

W reacts faster with dilute hydrochloric acid than Y.

The hydroxide of X is less soluble in water than the hydroxide of W.

The sulfate of Y is the most soluble of the sulfates of W, X, Y and Z.

The carbonate of Z decomposes more slowly than the carbonate of W at the same temperature.

What is the M_r of the nitrate of W?

A 148.3 **B** 164.1 **C** 211.6 **D** 261.3

25 Polymer J has repeat unit $–[CH(C_2H_5)CH(CH_3)]–$.

Polymer K has repeat unit $–[CH(CH_3)CH(CH_3)]–$.

Which row is correct?

	monomer from which polymer J is produced	monomer from which polymer K is produced
A	pent-1-ene	but-1-ene
B	pent-1-ene	but-2-ene
C	pent-2-ene	but-1-ene
D	pent-2-ene	but-2-ene

26 0.200 mol ethanenitrile reacts with an excess of dilute sodium hydroxide. The reaction produces organic compound S and ammonia gas only.

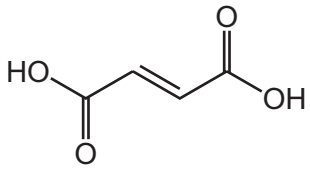
The reaction has an 80.0% yield.

Which mass of S is produced?

A 9.60 g **B** 13.1 g **C** 15.4 g **D** 16.4 g

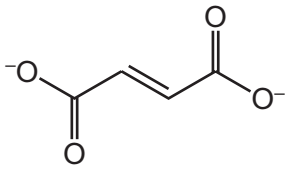
27 The structure of butenedioic acid is shown.

butenedioic acid



When an excess of NaOH(aq) is added to butenedioic acid, which organic product is formed?

A



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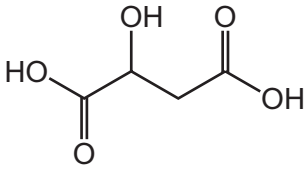

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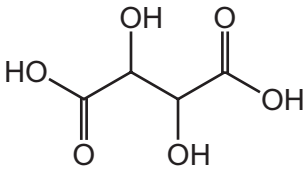

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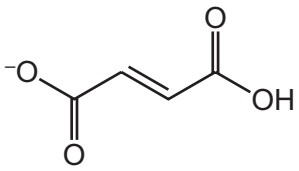
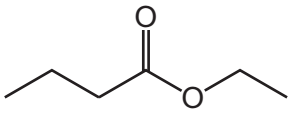
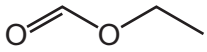
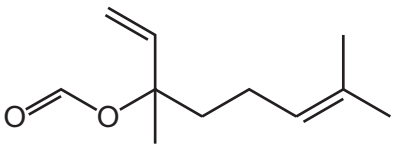
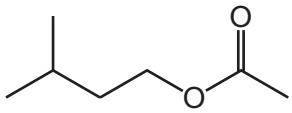


28 Four organic compounds are shown.









A mixture of these four compounds reacts with hot aqueous sodium hydroxide.

The products of this reaction are acidified.

Which organic acids are present in the products?

A butanoic, ethanoic and propanoic acids

B butanoic, ethanoic and methanoic acids

C ethanoic, 3-methylbutanoic and propanoic acids

D methanoic, 3-methylbutanoic and propanoic acids

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29 Ethanal reacts with KCN dissolved in liquid HCN. This reaction involves the formation of an intermediate.

Which statement is correct?

A HCN does **not** have any lone pairs of electrons and so CN^- is the catalyst.

B The $O^{\delta-}$ of C=O attacks K^+ in a nucleophilic attack.

C A proton from HCN is transferred to the intermediate.

D The reaction is a nucleophilic substitution.

30 What is formed when propanone is heated under reflux with a solution of $NaBH_4$?

A propan-2-ol

B propan-1-ol

C propanal

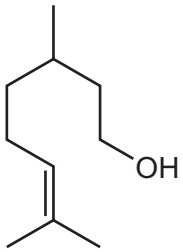
D propane

31 The compounds shown are all produced by plants.

Each compound is warmed with acidified $K_2Cr_2O_7$.

Which compound will give a different observation to the other three?

A



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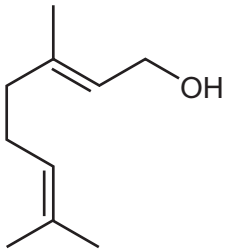

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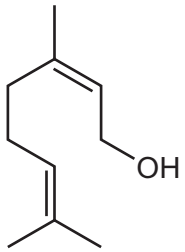

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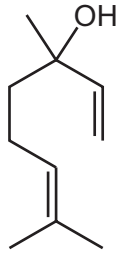

32 Structural and stereoisomerism should be considered when answering this question.

How many alcohols with the molecular formula $C_5H_{12}O$ give a yellow precipitate with alkaline I_2(aq)?

A 2 **B** 3 **C** 4 **D** 5

33 An amine is produced in the following reaction.

$$C_2H_5I + 2NH_3 \rightarrow C_2H_5NH_2 + NH_4I$$

What is the mechanism?

A electrophilic addition

B free-radical substitution

C nucleophilic addition

D nucleophilic substitution

34 The reaction shown can be used to lengthen a carbon chain.

$$CH_3CH_2CH_2Br + CN^- \rightarrow CH_3CH_2CH_2CN + Br^-$$

Which row shows the correct reagent and conditions for this reaction?

	reagent	conditions
A	KCN	heat under reflux in dilute sulfuric acid
B	KCN	heat under reflux in ethanol
C	HCN	heat under reflux in dilute sulfuric acid
D	HCN	heat under reflux in ethanol

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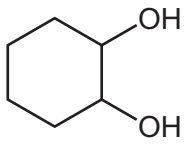
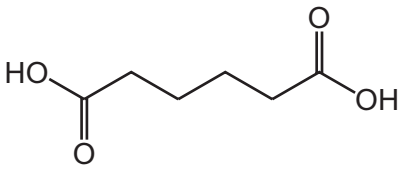
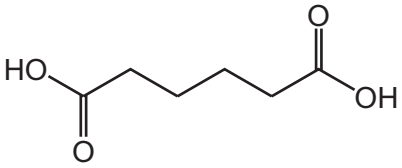
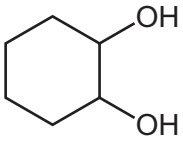
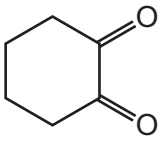
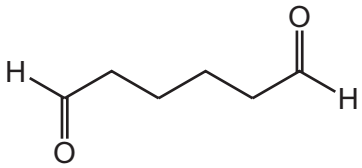
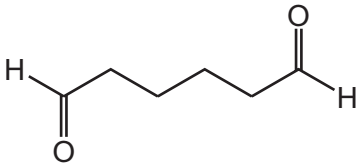
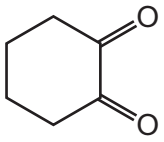
35 Which row shows the products formed when cyclohexene reacts with acidified $KMnO_4$ under different conditions?

	product formed on reaction with cold, dilute acidified $KMnO_4$	product formed on reaction with hot, concentrated acidified $KMnO_4$
A		
B		
C		
D		

36 Two steps in the free-radical substitution reaction between methane and chlorine are shown.

step 1 $CH_3\bullet + Cl_2 \rightarrow CH_3Cl + Cl\bullet$

step 2 $CH_3Cl + Cl\bullet \rightarrow CH_2Cl\bullet + HCl$

Which statement is correct?

A Step 1 is initiation and step 2 is propagation.

B Step 1 is propagation and step 2 is termination.

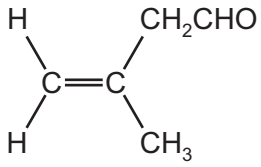
C Step 1 is initiation and step 2 is termination.

D Both steps are propagation.

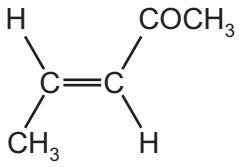
37 Compound P displays cis/trans isomerism and gives a red-brown precipitate with Fehling's solution.

What is P?

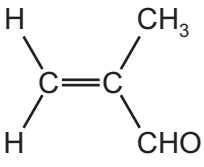
A



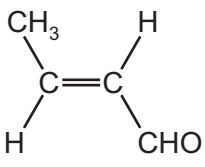
B



C

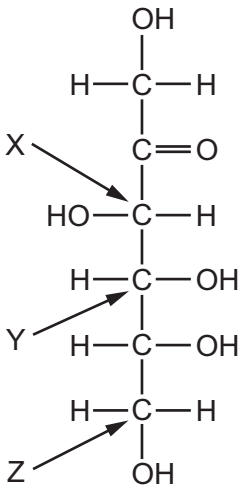


D



38 Fructose is a sugar with more than one chiral centre. The fructose molecule is shown with X, Y and Z indicating three carbon atoms.

fructose



Which carbon atoms are chiral centres?

A X, Y and Z **B** X and Y only **C** X only **D** Y only

39 An organic compound, T, contains only one functional group.

Compound T has the empirical formula C_2H_4O.

Some functional groups are listed.

1 alcohol

2 aldehyde

3 ester

4 ketone

Which functional groups are possible for compound T?

A 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 2 and 4

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40 The mass spectrum of an organic compound has the following features.

- The molecular ion peak is at $m/e = 60$. This peak has a relative intensity = 100.
- There is a peak at $m/e = 61$, which has a relative intensity = 2.2.
- There is **no** fragment peak at $m/e = 17$.
- There is a fragment peak at $m/e = 31$.

What could be the identity of the organic compound?

A methyl methanoate, $HCOOCH_3$

B methoxyethane, $CH_3OC_2H_5$

C ethanoic acid, CH_3COOH

D propan-2-ol, $CH_3CH(OH)CH_3$

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Important values, constants and standards

molar gas constant	$R = 8.31\,\text{J K}^{-1}\,\text{mol}^{-1}$
Faraday constant	$F = 9.65 \times 10^{4}\,\text{C mol}^{-1}$
Avogadro constant	$L = 6.022 \times 10^{23}\,\text{mol}^{-1}$
electronic charge	$e = -1.60 \times 10^{-19}\,\text{C}$
molar volume of gas	$V_m = 22.4\,\text{dm}^3\,\text{mol}^{-1}$ at s.t.p. (101 kPa and 273 K) $V_m = 24.0\,\text{dm}^3\,\text{mol}^{-1}$ at room conditions
ionic product of water	$K_w = 1.00 \times 10^{-14}\,\text{mol}^2\,\text{dm}^{-6}$ (at 298 K (25 °C))
specific heat capacity of water	$c = 4.18\,\text{kJ kg}^{-1}\,\text{K}^{-1}$ ($4.18\,\text{J g}^{-1}\,\text{K}^{-1}$)

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The Periodic Table of Elements

Group																	
1	2											13	14	15	16	17	18
							1 H hydrogen 1.0										2 He helium 4.0
3 Li lithium 6.9	4 Be beryllium 9.0											5 B boron 10.8	6 C carbon 12.0	7 N nitrogen 14.0	8 O oxygen 16.0	9 F fluorine 19.0	10 Ne neon 20.2
11 Na sodium 23.0	12 Mg magnesium 24.3	3	4	5	6	7	8	9	10	11	12	13 Al aluminium 27.0	14 Si silicon 28.1	15 P phosphorus 31.0	16 S sulfur 32.1	17 Cl chlorine 35.5	18 Ar argon 39.9
19 K potassium 39.1	20 Ca calcium 40.1	21 Sc scandium 45.0	22 Ti titanium 47.9	23 V vanadium 50.9	24 Cr chromium 52.0	25 Mn manganese 54.9	26 Fe iron 55.8	27 Co cobalt 58.9	28 Ni nickel 58.7	29 Cu copper 63.5	30 Zn zinc 65.4	31 Ga gallium 69.7	32 Ge germanium 72.6	33 As arsenic 74.9	34 Se selenium 79.0	35 Br bromine 79.9	36 Kr krypton 83.8
37 Rb rubidium 85.5	38 Sr strontium 87.6	39 Y yttrium 88.9	40 Zr zirconium 91.2	41 Nb niobium 92.9	42 Mo molybdenum 95.9	43 Tc technetium –	44 Ru ruthenium 101.1	45 Rh rhodium 102.9	46 Pd palladium 106.4	47 Ag silver 107.9	48 Cd cadmium 112.4	49 In indium 114.8	50 Sn tin 118.7	51 Sb antimony 121.8	52 Te tellurium 127.6	53 I iodine 126.9	54 Xe xenon 131.3
55 Cs caesium 132.9	56 Ba barium 137.3	57–71 lanthanoids	72 Hf hafnium 178.5	73 Ta tantalum 180.9	74 W tungsten 183.8	75 Re rhenium 186.2	76 Os osmium 190.2	77 Ir iridium 192.2	78 Pt platinum 195.1	79 Au gold 197.0	80 Hg mercury 200.6	81 Tl thallium 204.4	82 Pb lead 207.2	83 Bi bismuth 209.0	84 Po polonium –	85 At astatine –	86 Rn radon –
87 Fr francium –	88 Ra radium –	89–103 actinoids	104 Rf rutherfordium –	105 Db dubnium –	106 Sg seaborgium –	107 Bh bohrium –	108 Hs hassium –	109 Mt meitnerium –	110 Ds darmstadtium –	111 Rg roentgenium –	112 Cn copernicium –	113 Nh nihonium –	114 Fl flerovium –	115 Mc moscovium –	116 Lv livermorium –	117 Ts tennessine –	118 Og oganesson –

Key

atomic number
atomic symbol
name
relative atomic mass

lanthanoids	57 La lanthanum 138.9	58 Ce cerium 140.1	59 Pr praseodymium 140.9	60 Nd neodymium 144.2	61 Pm promethium –	62 Sm samarium 150.4	63 Eu europium 152.0	64 Gd gadolinium 157.3	65 Tb terbium 158.9	66 Dy dysprosium 162.5	67 Ho holmium 164.9	68 Er erbium 167.3	69 Tm thulium 168.9	70 Yb ytterbium 173.1	71 Lu lutetium 175.0
actinoids	89 Ac actinium –	90 Th thorium 232.0	91 Pa protactinium 231.0	92 U uranium 238.0	93 Np neptunium –	94 Pu plutonium –	95 Am americium –	96 Cm curium –	97 Bk berkelium –	98 Cf californium –	99 Es einsteinium –	100 Fm fermium –	101 Md mendelevium –	102 No nobelium –	103 Lr lawrencium –